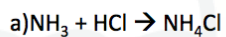
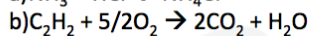
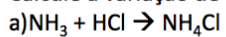


Termoquímica - Entalpia na Reação

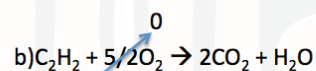
Calor de formação: $\Delta H = \Delta H_p - \Delta H_R$

→ Substância Simples no estado padrão → $\Delta H = 0$

Calcule a variação de entalpia para as reações abaixo:



$$\Delta H = -315 + 138 = -177 \text{ KJ/mol}$$



$$+226 \quad \quad \quad \rightarrow -1071$$

$$\Delta H = -1071 - 226 = -1297 \text{ KJ/mol}$$

| Substância | ΔH (KJ/mol) |
|------------------------|---------------------|
| NH_3 | -46 |
| HCl | -92 |
| NH_4Cl | -315 |
| C_2H_2 | +226 |
| CO_2 | -393 |
| H_2O | -285 |